Assessment on d block elements.

1. Why elements do were classify under a specific group called “d”?
2. Give the symbols of 3d elements.
3. 27 is the atomic number of a 3d element called “M”.
   i) Give the electronic configuration of the ground state atom “M”.
   ii) Suppose M makes a divalent caution, give the electronic configuration of M2+.
   iii) How many unpaired electrons does M2+ ion bare
4. What are transition element ?
   i) why do certain d elements are not considered as transition?
   Ii) Name them
5. Say whether the followings are true or false.
   a. S block elements make cautions preferably than d block ( )
   b. S block elements are denser than d block elements ( )
   c. Metallic lattice of d block elements are much stronger ( )
   d. All the industrial catalysts are d block metals ( )
   e. Cr,Mn,Fe,Zn are called transition metals while Sc is not( )
   f. Coordination number of Cr is 5 in$[Cr(H_2O)_5\,OH]^{2+}$ ( )
   g. Ligands are electro donars while d block cations behave as Lewis acides ( )
   h. Geometry Is angular in $[Ag(NH_3)_2$ ( )
   i. $CuCl_4$ is tetrahedral while [Cu(NH3)4]2+ is square planar although the coordination number is 4 ( )

Name the following complex ions and compounds.

i)$[Ni(NH_3)_6]^{2+}$

ii)$K_4[Fe(CN)_6]$

iii)$KCa[Fe(CN)_6]$

iV) $Cu(NH_3)_4Br_2$

V)$Ag(NH_3)_2I$